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 $6.02 \times 10^{23} \text{ atoms mol}^{-1} \times 47.0 \text{ electrons atom}^{-1} = 2.62 \times 10^{24}$ (b) #
electrons added = $Q / e = 1.00 \times 10^{23} \text{ C} / 1.60 \times 10^{-19} \text{ C electron}^{-1} = 6.25$
 $\times 10^{15}$ or 2.38 electrons for every 10^9 already present 23.2 (a) $F_e = k_e q_1 q_2 / r^2 = 8.99 \times 10^9 \text{ N m}^2/\text{C}^2 \times (1.60 \times 10^{-19} \text{ C})^2 / (3.80 \times 10^{-10} \text{ m})^2 = 1.59 \times 10^{-9} \text{ N}$ (repulsion) (b) $F_g = G m_1 m_2$

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